



This is “Appendix D: Dissociation Constants and pK<sub>b</sub> Values for Bases at 25°C”, appendix 4 from the book [Principles of General Chemistry \(index.html\)](#) (v. 1.0).

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## Chapter 28

### Appendix D: Dissociation Constants and $pK_b$ Values for Bases at 25°C

Name	Formula	$K_b$	$pK_b$
Ammonia	$NH_3$	$1.8 \times 10^{-5}$	4.75
Aniline	$C_6H_5NH_2$	$7.4 \times 10^{-10}$	9.13
<i>n</i> -Butylamine	$C_4H_9NH_2$	$4.0 \times 10^{-4}$	3.40
<i>sec</i> -Butylamine	$(CH_3)_2CHCH_2NH_2$	$3.6 \times 10^{-4}$	3.44
<i>tert</i> -Butylamine	$(CH_3)_3CNH_2$	$4.8 \times 10^{-4}$	3.32
Dimethylamine	$(CH_3)_2NH$	$5.4 \times 10^{-4}$	3.27
Ethylamine	$C_2H_5NH_2$	$4.5 \times 10^{-4}$	3.35
Hydrazine	$N_2H_4$	$1.3 \times 10^{-6}$	5.9
Hydroxylamine	$NH_2OH$	$8.7 \times 10^{-9}$	8.06
Methylamine	$CH_3NH_2$	$4.6 \times 10^{-4}$	3.34
Propylamine	$C_3H_7NH_2$	$3.5 \times 10^{-4}$	3.46
Pyridine	$C_5H_5N$	$1.7 \times 10^{-9}$	8.77
Trimethylamine	$(CH_3)_3N$	$6.3 \times 10^{-5}$	4.20

Source of data: *CRC Handbook of Chemistry and Physics*, 84th Edition (2004).